

SIR C.R.REDDY COLLEGE FOR WOMEN, ELURU
CURRICULUM LECTURER WISE 2019-2020

ANNUAL CURRICULAR					PLAN (Year)				CO-CURRICULAR ACTIVITY			
NAMR OF THE LECTURER <i>B. Tulani Kotewaribai</i>					CLASS : <i>BSc mcs-1</i> Semester : <i>1</i>				Paper : <i>chemistry-I</i>			
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date
Novem	3 rd	4	Syllabus Dictation (1h) Basics of chemistry (2h)	—	—	—	—	—	—	—	—	—
			unit-1 Chemistry of pblock elements Group-13:	Electronic Configuration of Group-13	—	—	—	—	—	—	—	—
			Preparation and structure of Diborane.	—	—	—	—	—	—	—	—	—
	4 th	5	Preparation and structure of Borazine Group-14:	—	—	—	—	—	—	—	—	—
			Preparation, classification and uses of silicones.	—	—	—	—	—	—	—	—	—
			Preparation and structures of phospho nitric halides ($PNCl_2$) _n ; n=3,4	—	—	—	—	—	—	—	—	—

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SIR C.R.REDDY COLLEGE FOR WOMEN, ELURU
CURRICULUM LECTURER WISE 2019-2020

ANNUAL CURRICULAR					PLAN (Year)				Inorganic and physical chemistry				
NAMR OF THE LECTURER : B. Tulani Koteswari bai					CLASS : B.Sc. MCCS-4		Semester : I		Paper : Chemistry-I				
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
December	1st	5	Group-16: Oxides and oxyacids of Sulphur (structures only)	—	—	—	—	—	—	—	—	—	—
			Group-17: Pseudo-halogens, structure of inter halogen compounds.	—	—	—	—	—	—	—	—	—	—
			unit-II 1. Chemistry of d-block elements: characteristics of	—	—	—	—	—	—	—	—	—	—
			d-block elements with special reference to electronic configuration.	—	—	—	—	—	—	—	—	—	—
	2nd	5	Variable valence, magnetic properties, catalytic properties and	—	—	—	—	—	—	—	—	—	—
			ability to form complexes. Stability of various oxidation states.	—	—	—	—	—	—	—	—	—	—

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SIR C.R.REDDY COLLEGE FOR WOMEN, ELURU
CURRICULUM LECTURER WISE 2019 - 2020

ANNUAL CURRICULAR					PLAN (Year) <i>Inorganic and Physical Chemistry</i>								
NAME OF THE LECTURER : <i>B. Tulani Koteswari bai</i>					CLASS : <i>BSc MCCs-1</i>			Semester : <i>I</i>		Paper : <i>Chemistry-I</i>			
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
	<i>3rd</i>	<i>5</i>	<i>2) Chemistry of f block elements: chemistry of lanthanides - electronic</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>
			<i>Structure, oxidation states, lanthanide contraction, consequences of lanthanide</i>	<i>Separation of Lanthanides by Ion-exchange method</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>
			<i>Contraction, magnetic properties. Chemistry of actinides - electronic</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>
			<i>Configuration, oxidation states.</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>
	<i>4th</i>	<i>4</i>	<i>Actinide Contraction Comparison of Lanthanides and actinides.</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>
			<i>3) Theories of Bonding in metals: Valence bond theory and Free electron Theory</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>	<i>—</i>

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SIR C.R.REDDY COLLEGE FOR WOMEN, ELURU
CURRICULUM LECTURER WISE 2019-2020

ANNUAL CURRICULAR					PLAN (Year) Inorganic and physical chemistry								
NAMR OF THE LECTURER : B. Tulani Koteswari bai					CLASS : I BSc MCCS-1 Semester : 2				Paper : Chemistry 3				
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO- CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
			Explanation of thermal and electrical conductivity of metals based on these theories.	—	—	—	—	—	—	—	—	—	—
January	1st	5	Band theory-formation of bands, explanation of conductors.	—	—	—	—	—	—	—	—	—	—
			Semi conductors and insulators unit-IV 1. Gaseous state.	Boyle's law, Charles law & Avagadro's law.	—	—	—	—	—	—	—	—	—
			Vander waal's equation of state, Andrews isotherms of CO ₂ , Continuity of state.	Derivation of Real gases from Gas laws.	—	—	—	—	—	—	—	—	—
			Critical phenomena. Relationship between critical constants and Vander waal's constants.	—	—	—	—	—	—	—	—	—	—
	3rd	5	Law of corresponding states. Joule-Thomson effect. Inversion temp.	—	—	—	—	—	—	—	—	—	—

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CURRICULUM LECTURER WISE 2019-2020

ANNUAL CURRICULAR					PLAN (Year) <i>Inorganic and Physical chemistry</i>								
NAMR OF THE LECTURER : <i>B. Tulani Koteswari Bai</i>					CLASS : <i>B.Sc MCC-1</i>			Semester : <i>I</i>		Paper : <i>Chemistry-I</i>			
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
			<i>2. Liquid state: Liquid crystals, mesomorphic state</i>	—	—	—	—	—	—	—	—	—	—
			<i>Differences between liquid crystal and solid/liquid. classification of</i>	—	—	—	—	—	—	—	—	—	—
			<i>liquid crystals into Smectic and Nematic.</i>	—	—	—	—	—	—	—	—	—	—
	<i>4th</i>	<i>4</i>	<i>Application of liquid crystals as LCD devices.</i>	—	—	—	—	—	—	—	—	—	—
			<i>Unit-III Solid state. Symmetry in crystals. Law of</i>	—	—	—	—	—	—	—	—	—	—
			<i>constancy of interfacial angles. The law of rationality of indices</i>	—	—	—	—	—	—	—	—	—	—

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CURRICULUM LECTURER WISE 2019-2020

ANNUAL CURRICULAR					PLAN (Year) <i>Inorganic and Physical Chemistry</i>								
NAMR OF THE LECTURER : <i>B. Tulani Kotevaran Bai</i>					CLASS: <i>B.Sc. MCC-1</i> Semester: <i>T</i>				Paper: <i>chemistry-I</i>				
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
			<i>The law of Symmetry. Miller indices.</i>	—	<i>Plane of Symmetry Axis of Symmetry models shown</i>	—	—	—	—				
<i>February</i>	<i>1st</i>	<i>5</i>	<i>Definition of lattice point, space lattice, unit-cell. Bravais</i>	—	—	—	—	—	—				
			<i>lattices and crystal systems. X-ray diffraction and</i>	—	—	—	—	—	—				
			<i>Crystal structure. Bragg's law Powder method. Defects in</i>	<i>Structure of NaCl is to be explained.</i>	—	—	—	—	—				
			<i>Crystals. Stoichiometric and non-stoichiometric defects.</i>	—	—	—	—	—	—				
	<i>2nd</i>	<i>5</i>	<i>und=V 1. Solutions: Azeotropes HCl-H₂O system and ethanol-water system</i>	<i>Definitions of terms used to represent concentration</i>	—	—	—	—	—				

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ANNUAL CURRICULAR					PLAN (Year)							
NAME OF THE LECTURER : B. Tulani Kotewani bai					CLASS : B.Sc M.C.S.T Semester : 7							
					CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY			
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date
			Partially miscible liquids - phenol water system. Critical solution temperature (CST)	—	—	—	—	—	—	—	—	—
			Effect of impurity on Consolute temperature	—	—	—	—	—	—	—	—	—
			Immiscible liquids and steam distillation.	—	—	—	—	—	—	—	—	—
	3rd	5	Nernst distribution law. Calculation of the partition coefficient.	—	—	—	—	—	—	—	—	—
			Applications of distribution law.	—	—	—	—	—	—	—	—	—
			2) Ionic equilibrium	—	—	—	—	—	—	—	—	—
			Ionic product, common ion effect, solubility and solubility product calculations based on solubility product.	—	—	—	—	—	—	—	—	—
	4th	4	3) Dilute Solutions :- Colligative properties - RLVP, osmotic pressure.	—	—	—	—	—	—	—	—	—
			Elevation in boiling point and depression in freezing point.	—	—	—	—	—	—	—	—	—

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ANNUAL CURRICULAR					PLAN (Year) Inorganic and Physical Chemistry								
NAMR OF THE LECTURER					CLASS : BSc mcs-I Semester : I Paper : Chemistry-I								
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CO CURRICULAR ACTIVITY				GO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
			Experimental methods for the determination of molar mass of a non-volatile solute using osmotic pressure.	—	—	—	—	—	—	—	—	—	—
March	1st	5	Experimental method for the determination of Elevation in boiling point and depression in freezing point. Abnormal colligative properties. Van't Hoff factor.	—	—	—	—	—	—	—	—	—	—
	2nd	5	—	—	—	—	—	—	—	—	—	—	—
			—	—	—	—	—	—	—	—	—	—	—
	3rd	5	—	—	—	—	—	—	—	—	—	—	—
			—	—	—	—	—	—	—	—	—	—	—

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ANNUAL CURRICULAR					PLAN (Year)								
NAMR OF THE LECTURER <u>P. RAMYA KRISHNA</u>					CLASS : <u>II BFC</u>			Semester : <u>III</u>			Paper : <u>organic and spectroscopy</u>		
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO- CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
Novem ber	2 nd	2	Syllabus- dictation preparation and reactions of Alkyl Halides	-	-	-	-	-	-	-	-	-	-
			SN ¹ and SN ² reactions, williamson's synthesis.	-	-	-	-	-	-	-	-	-	-
	3 rd	5	Reactivities of halogen on alkyl aryl, allyl, vinyl aryl alkyl halides	-	-	-	-	-	-	-	-	-	-
	4 th	5	Preparation methods of alcohols and phenols, Acidic strength of phenols	-	-	-	-	-	-	-	-	-	-
Decem ber	1 st	5	Reimer - Tiemann, Kolbe - Schmidt, pinacole - pinacolone claisen Schmidt Fries, acetic acid synthesis	-	-	-	-	-	-	-	-	-	-

P. Ramya
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ANNUAL CURRICULAR					PLAN (Year)								
NAMR OF THE LECTURER <u>P. RAMYA KRISHNA</u>					CLASS : <u>TI 2FC</u>		Semester : <u>III</u>		Paper : <u>organic and spectroscopy</u>				
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO- CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
	2nd	5	chemical reactions of Alcohols and phenols, Introduction to carbonyl compounds, and I mid examinations	will be shown preparation of phenol and their reactivity	-	4	yes	-	-	-	-	-	-
					-	-	-	-	-	-	-	-	-
					-	-	-	-	-	-	-	-	-
	3rd	5	Preparation methods of aldehydes & ketones, reactivity of carbonyl group, Nucleophilic addition reactions Aldol condensation Benzoin condensation		-	-	-	-	-	-	-	-	-
					-	-	-	-	-	-	-	-	-
					-	-	-	-	-	-	-	-	-
	4th	5	Cannizzaro reaction Wittig, Perkin's reaction, Clemmenson's reduction, Baeyer-Villiger, Claisen condensation, Haloform reaction		-	-	-	-	-	-	-	-	-
					-	-	-	-	-	-	-	-	-
					-	-	-	-	-	-	-	-	-

Signature of the Lecturer P. Ramya

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ANNUAL CURRICULAR					PLAN (Year)								
NAME OF THE LECTURER: P. RAMYA KRISHNA					CLASS: II 2FC			Semester: III		Paper: organic and spectroscopy			
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
Jan	1st	5	Reduction reactions preparation and chemical	-	-	-	-	-	-	-	-	-	-
			properties of Aceto acetic ester and malonic ester	-	-	-	-	-	-	-	-	-	-
	2nd	5	Pongal Holidays	-	-	-	-	-	-	-	-	-	-
	3rd	5	Introduction to Carboxylic acids preparation and chemical reactions of carboxylic acids,	-	-	-	-	-	-	-	-	-	-
			Reformatsky, HVZ reaction, esterification	Hydrolysis of esters - demonstration	experiment on ester hydrolysis	01	yes	-	-	-	-	-	-
	4th	5	hydrolysis of esters	-	-	-	-	-	-	-	-	-	-
			Introduction to molecular spectroscopy and types of molecular spectra	will be shown PPT presentation of molecular spectra	PPT presentation	02	yes	-	-	-	-	-	-

Signature of the Lecturer: P. Ramya

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SIR C.R.REDDY COLLEGE FOR WOMEN, ELURU
CURRICULUM LECTURER WISE 2019-2020

ANNUAL CURRICULAR					PLAN (Year)											
NAME OF THE LECTURER: P. RAMYA KRISHNA					CLASS: II 2FC				Semester: III				Paper: organic and spectroscopy			
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY							
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date				
February	1st	5	Vibrational spectroscopy-spectral lines Intensity of spectral lines	-	debate	01	yes		debate	01	yes					
			rotational spectroscopy - harmonic and Anharmonic oscillator	-	-	-	-	-	-	-	-	-				
	2nd	5	Types of electronic transitions Beer Lambert's law, effect of conjugation,	-	seminar	01	yes		seminar debate	01	yes					
			chromophore and Auxochrome Intensity shifts of absorption	will be shown PPT presentation of absorption	PPT presentation of absorption	01	yes		-	-	-	-				
	3rd	5	principle of NMR, equivalent non equivalent protons spin-spin coupling, coupling constant,	intensity shift	shifts	-	-	-	-	-	-	-				
			Applications of NMR Spectroscopy	-	-	-	-	-	-	-	-	-				

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ANNUAL CURRICULAR					PLAN (Year)								
NAME OF THE LECTURER: P. RAMYA KRISHNA					CLASS: II 2FC			Semester: III		Paper: organic and Spectroscopy			
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
	4th	4	Applications of NMR, chemical shift	will be shown PPT presentation of principle of NMR	presentation of NMR principle PPT	01	yes	-	-	-	-	-	-
MAR CH	1st	5	pre-final examinations	-	Quiz	01	yes	-	Quiz	01	yes	-	-
	2nd	5	Applications of woodward-Hoffmann rules to carbonyl compounds and α, β unsaturated acids	-	-	-	-	-	-	-	-	-	-
	3rd	5	Functional group and fingerprint region, characteristic absorption bands of various,	-	Group Discussion	01	yes	-	Group Discussion	01	yes	-	-
			Interpretation of IR Spectra	Remedial class	-	-	-	-	-	-	-	-	-

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SIR C.R.REDDY COLLEGE FOR WOMEN, ELURU

CURRICULUM LECTURER WISE 2021 - 2022

ANNUAL CURRICULAR					PLAN (Year) <i>Inorganic, organic & physical chemistry</i>							
NAME OF THE LECTURER <i>S.R. Tejaswi</i>					CLASS : <i>III BSc (MCC)</i>				Semester : <i>VA</i>		Paper : <i>VA</i>	
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY			
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date
<i>Oct</i>	<i>1st</i>	<i>2</i>	<i>Syllabus dictation. Unit-I Coordination chemistry definitions of terms used in coordination chemistry.</i>	<i>Types of chemical bonds- Ionic, covalent & co-ordinate covalent bonds.</i>	-	-	-	-	-	-	-	-
					-	-	-	-	-	-	-	-
	<i>2nd</i>	<i>4</i>	<i>IUPAC Nomenclature bonding theories - Review of Werner's theory & Sidwick's concept of co-ordination. VBT</i>	<i>Double salts & co-ordination compounds.</i>	-	-	-	-	-	-	-	-
					-	-	-	-	-	-	-	-
	<i>3rd</i>	<i>3</i>	<i>Geometries of C.N's & tetrahedral & square planar &</i>		-	-	-	-	-	-	-	-
				<i>6-octahedral & its limitations, CFT-splitting of d-orbitals</i>		<i>Remedial class</i>	<i>01</i>	<i>Yes</i>	-	-	-	-

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SIR C.R.REDDY COLLEGE FOR WOMEN, ELURU
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ANNUAL CURRICULAR					PLAN (Year)								
NAME OF THE LECTURER: <u>S.R. Tejasa</u>					CLASS: <u>III B.Sc (MCG)</u> Semester: <u>V</u>				Paper: <u>NA</u> <u>Inorganic, organic & physical chem</u>				
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
			in octahedral, tetrahedral & square planar complexes. low spin & high spin complexes	-	-	-	-	-	-	-	-	-	-
	4 th	4	Factors affecting CFSE, merits & demerits of CFT, Isomerism in	-	-	-	-	-	-	-	-	-	-
			co-ordination comp's. Structural isomerism & Stereo isomerism,	-	Remedial class	01	Yes	-	-	-	-	-	-
			Stereo isomerism of complexes with 4 & 6 coordination numbers.	-	-	-	-	-	-	-	-	-	-
			Unit III 1. Spectral & Magnetic properties of metal complexes.	Spectra of $[Ti(H_2O)_6]^{3+}$ ion.	will be shown PPT	01	Yes	-	-	-	-	-	-
			Types of magnetic behaviour	-	-	-	-	-	-	-	-	-	-

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CURRICULUM LECTURER WISE 2021- 2022

ANNUAL CURRICULAR					PLAN (Year)				Inorganic, organic & physical chemistry				
NAME OF THE LECTURER <u>S.R. Tejaswi</u>					CLASS : <u>III BSc (MCG)</u> Semester : <u>V</u>				Paper : <u>VA</u>				
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
Nov	1st	4	Spins only formula, calculation of magnetic moments, experimental determination of magnetic susceptibility	-	Remedial	01	Yes	-	-	-	-	-	-
			- Gouy method. 2. Stability of metal complexes.	-	will be shown Magnetic susceptibility (PPT)	01	Yes	-	-	-	-	-	-
			Thermodynamic stability & kinetic stability	-	-	-	-	-	-	-	-	-	-
	2nd	4	Factors affecting the stability of metal complexes. chelate effect,	-	Quiz	01	Yes	-	-	-	-	-	-
			determination of composition of complex by Job's method & mole-ratio method.	-	-	-	-	-	-	-	-	-	-
			Unit-III Nitrohydro Carbons.	-	-	-	-	-	-	-	-	-	-

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S.R. Tejaswi

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B. Anuradha

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SIR C.R.REDDY COLLEGE FOR WOMEN, ELURU
CURRICULUM LECTURER WISE 2021 - 2022

ANNUAL CURRICULAR					PLAN (Year)				CO-CURRICULAR ACTIVITY			
NAME OF THE LECTURER <u>S.R. Tejaswi</u>					CLASS : <u>III BSc (MCG)</u> Semester : <u>V</u> Paper : <u>VA</u>				<u>Inorganic, organic & physical chemistry</u>			
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date
			Nomenclature & classification of Nitro hydrocarbons, structure Tautomerism	-	-	-	-	-	-	-	-	-
			of nitroalkanes leading to acid & keto form.	-	-	-	-	-	-	-	-	-
	3 rd	4	Preparation of Nitroalkanes, reactivity - halogenation, reaction with HNO_2 , Nef rea ⁿ & Mannich rea ⁿ leading to Michael add ⁿ & reduction.	-	Remedial	1	Yes	-	-	-	-	-
			unit-IV	-	-	-	-	-	-	-	-	-
			Nitrogen compounds: Amines (Aliphatic) Nomenclature,	-	-	-	-	-	-	-	-	-
			classification into 1 ^o , 2 ^o & 3 ^o amines & 4 ^o ammonium comp's	-	-	-	-	-	-	-	-	-

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CURRICULUM LECTURER WISE 2021 - 2022

ANNUAL CURRICULAR					PLAN (Year)				Inorganic, organic & physical chemistry				
NAMR OF THE LECTURER S.R. Tejaswi					CLASS : III BSc (MCC) Semester : V				Paper : IA				
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
			ammonium comp's Preparative methods 1. Ammonolysis of alkyl halides.	-	-	-	-	-	-	-	-	-	-
			2. Gabriel synthesis 3. Hoffmann's bromamide re ⁿ (mech), Reduction of	-	-	-	-	-	-	-	-	-	-
	4 th	4	Amides & Schmidt re ⁿ , chemical Prop a) alkylation b) Acylation.	-	Remedial	01	Yes	-	-	-	-	-	-
			c) Carbyl amine re ⁿ & d) Hinsberg separation e) with nitrous acid	-	-	-	-	-	-	-	-	-	-
Dec	1 st	4	oxidation of 3 ^o amines, Nomenclature, classification of Aromatic Amines.	-	-	-	-	-	-	-	-	-	-
			Prep of Aromatic amines - Ammonolysis of aryl halides,	-	-	-	-	-	-	-	-	-	-

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ANNUAL CURRICULAR					PLAN (Year)				CO-CURRICULAR ACTIVITY			
NAME OF THE LECTURER <u>S.R. Tejaswi</u>					CLASS <u>III B.Sc (MCG)</u> Semester: <u>V</u> Paper: <u>VA</u>				Inorganic, organic & physical chemistry			
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			Hoffmann bromamide react ⁿ & Schmidt react ⁿ . chemical prop of aromatic amines	-	-	-	-	-	-	-	-	-
			- Alkylation, Acylation Gabyl amine react ⁿ , Diazotization, Oxidation, electrophilic	-	-	-	-	-	-	-	-	-
			substitution & Bromination Nitration	-	-	-	-	-	-	-	-	-
			Comparative Basic strength of ammonia, methyl amine, dimethyl amine, tri	-	-	-	-	-	-	-	-	-
			methyl amine & aniline.	-	-	-	-	-	-	-	-	-
	<u>end</u>	<u>4</u>	Comparative Basic strength of aniline, N-methyl aniline & N,N-dimethyl aniline	-	-	-	-	-	-	-	-	-

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CURRICULUM LECTURER WISE 2021- 2022

ANNUAL CURRICULAR					PLAN (Year)				Inorganic, organic & physical chemistry				
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MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
			(in aq. medium) steric & substituent effects. Unit-V Thermodynamics	-	-	-	-	-	-	-	-	-	-
	3 rd	4	First law of Thermodynamics - statement, Def of internal energy & enthalpy, heat capacities & their relation ship. Joule-Thomson effect - coefficient	Definitions of terms used in Thermodynamics.	Group Discussion	1	Yes	-	-	-	-	-	-
			Calculation of w , for the expansion of perfect gas under isothermal & adiabatic conditions for reversible process, state function.	Derivation of $PV^{\gamma} = \text{constant}$ $TV^{\gamma-1} = \text{constant}$	-	-	-	-	-	-	-	-	-
	4 th	4	Temp dependance of enthalpy of formation - kirchoff eq ⁿ .	-	Debate	1	Yes	-	-	-	-	-	-

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Jan	1st	4	2nd law of thermodynamics, 2nd statement of the law. Carnot cycle & its efficiency. Carnot theorem.	-	Seminar	01	Yes	-	-	-	-	-
			concept of entropy, entropy as a state function.	-	-	-	-	-	-	-	-	-
	2nd	2	Entropy changes in reversible & irreversible process. Entropy changes in spontaneous & equilibrium process.	-	-	-	-	-	-	-	-	-
	3rd	4	Mid exams preparation.	-	Revision	-	Yes	-	-	-	-	-
	4th	4	-	-	Remedial classes	04	Yes	-	-	-	-	-
Feb	1st	2	-	-	Remedial classes.	02	Yes	-	-	-	-	-

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CURRICULUM LECTURER WISE 2021- 2022

ANNUAL CURRICULAR					PLAN (Year)								
NAMR OF THE LECTURER V-RAJA RATESWARI					CLASS : III BSC		Semester : VB		Paper : VB				
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO- CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
Nov	3	1	mole ratio method, Thermodynamic & kinetic stability	-	-	-	-	-	-	-	-	-	-
	4	4	Nitro hydrocarbon Nomenclature, classification, Tautomerism, preparation, reactivity.	Definition of Nitrohydrocarbons with examples	-	-	-	-	-	-	-	-	-
	5	2	Nitrogen compound Aliphatic amines Preparation and reactivity	-	-	-	-	-	-	-	-	-	-
Dec	1	4	Aromatic amine Preparation and reactivity, Basic character	-	-	-	-	-	-	-	-	-	-
	2	4	steric effects and substituent effects Alkylation, Acylation, Carbylation, Hinsberg reaction	-	-	-	-	-	-	Group Discussion	1	yes	-
	3	4	Electrophilic substitution, Diazotization.	-	Preparation of azodye in lab	-	-	-	-	Debate	1	yes	-

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ANNUAL CURRICULAR					PLAN (Year)								
NAMR OF THE LECTURER					CLASS :	Semester :	Paper :			CO-CURRICULAR ACTIVITY			
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
Oct	2	4	Coordination Chemistry Introduction, IUPAC nomenclature	Definitions of the terms used in	-	-	-	-	-	-	-	-	
	3	2	Bonding theories - werner's theory, Sidgwick's theory	Coordination Chemistry	-	-	-	-	-	-	-	-	
	4	3	Valence bond theory, crystal field theory - octahedral,	-	-	-	-	-	-	-	-	-	
	5	4	tetrahedral and square planar complexes, Isomerism.	-	-	-	-	-	Quiz	1	yes	-	
Nov	1	4	Spectral and magnetic properties of metal complexes - Types, spin-only formula, Crystal field method.	-	-	-	-	-	-	-	-	-	
	2	2	stability of metal complexes - factors affecting the stability, Job's method.	kinetic stability and Thermo dynamic stability	-	-	-	-	-	-	-	-	

Signature of the Lecturer *V.R. Rajeswari*

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CURRICULUM LECTURER WISE 2021- 2022

ANNUAL CURRICULAR					PLAN (Year)								
NAMR OF THE LECTURER V-RAJA RATISWARI					CLASS : III BSC		Semester : IV		Paper : IV				
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO-CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
Dec	4	2	<u>Thermodynamics</u> First law statement Defn of internal energy.	Definitions in Thermodynamics.	-	-	-	-	-	-	-	-	-
	5	3	Defn of enthalpy Heat capacities Joule-Thomson effect	-	-	-	-	-	Group Discussion	1	yes	-	-
Jan	1	4	calculations of w for Isothermal expansion, adiabatic reversible expansion	-	-	-	-	-	student seminar	1	yes	-	-
	2		_____ Holidays	-	Sankranti Holidays				-	-	-	-	-
	3		_____	-	II Mid Exams				-	-	-	-	-
	4	3	Kirchoff's equation, Second law of thermodynamics.	-	-	-	-	-	-	-	-	-	-

Signature of the Lecturer **V.P. Raja**

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NAMR OF THE LECTURER V. RAJA RAJESWARI					CLASS : III BSC		Semester : IV		Paper : IV				
MONTH	WEEK	HOURS AVAILABLE	SYLLABUS/ TOPIC	Additional Input/Value Addition Provided/ Taught	CURRICULAR ACTIVITY				CO- CURRICULAR ACTIVITY				
					Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	Activity Conducted	Hours Allotted	Whether Conducted	If not Alternate Date	
Feb	1	4	II law states, Carnot cycle, Carnot theorem	-	-	-	-	-	-	-	-	-	-
	2	3	Definition of entropy, entropy changes in reversible process	-	-	-	-	-	-	-	-	-	-
	3	4	& Irreversible processes. Entropy change	-	-	-	-	-	-	-	-	-	-
	4	4	in spontaneous and equilibrium processes.	-	-	-	-	-	-	-	-	-	-
Mar	1	2											
	2	3											

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